

RESPONSE TO REVIEWERS

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Reviewer 2			
1.	Please mention the period of articles cited in this paper in the abstract.	We have included the period of articles cited in the abstract, covering articles published from 2002 to 2023.	1
2.	Please don't use pronouns in the text, such as we, our, etc.	We have made the necessary revisions to eliminate pronouns from the text, as per your request.	
3.	What do you mean by: "BPA has been classified as an endocrine-disrupting chemical (EDC)"?	BPA being classified as an endocrine-disrupting chemical (EDC) means it can interfere with the endocrine system, potentially causing health issues.	
4.	You should mention the drawbacks of each method presented in this review.	I have already revised the review to include the drawbacks of each method presented	6 - 32
5.	Explain the differences between "sieving mechanism" and "adsorption mechanism" based on membrane treatment.	The "sieving mechanism" relies on physical size-based separation, while the "adsorption mechanism" involves chemical interactions with contaminants. Sieving is size-selective, while adsorption is selective based on chemical affinity. Sieving is common in microfiltration and ultrafiltration, while adsorption is used for removing dissolved species and may involve membrane modification.	
6.	Based on your review, which membrane technology is more efficient for BPA removal under the same conditions?	The efficiency of membrane technology for BPA removal varies depending on factors like membrane properties, water pH, the presence of ions, and molecular weight cutoff. No single membrane technology appears universally superior, and the choice depends on specific water conditions and treatment goals.	
7.	What are the conditions of the wide range of half-lives stated by the following statement: "the half-lives of BPA and BPS range from 3 to 35 minutes"?	I apologize for any confusion, but the information about the conditions affecting the wide range of half-lives for BPA and BPS was already included in the original statement: "Lane et al. (2015) noted that in the presence of 1 mg/L of free chlorine, the half-lives of BPA and BPS range from 3 to 35 minutes, spanning a pH range of 6 to 11 and temperatures of 10-25°C."	20
8.	You stated that "Kinetic studies on BPA degradation by ferrate align well with a pseudo-second-order reaction law". Does that behavior always occur?	The behavior described in the statement, "Kinetic studies on BPA degradation by ferrate align well with a pseudo-second-order reaction law," suggests that in the specific studies mentioned, BPA	20

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		degradation by ferrate followed a pseudo-second-order reaction kinetics. However, it's important to note that reaction kinetics can vary depending on several factors, including reactant concentrations, reaction conditions, and the presence of other substances. Therefore, this behavior may not always occur under different experimental conditions or in all situations involving BPA degradation by ferrate. Kinetic behavior can be influenced by numerous factors, so it's essential to consider the specific context and conditions of any given study or application.	
9.	What is the limit of H ₂ O ₂ that you stated in the following statement: "Moreover, the introduction of high concentrations of hydrogen peroxide can contribute to the process of mineralization"?	The limit of H ₂ O ₂ concentration mentioned in the statement is not explicitly defined; however, the study by Zorzo et al. (2021) demonstrated that a concentration as high as 300 mg H ₂ O ₂ L ⁻¹ was effectively used in BPA wastewater treatment without any specific mention of an upper limit. The research showed that higher H ₂ O ₂ concentrations can accelerate BPA degradation and mineralization, but it also emphasized the importance of considering environmental sustainability. Therefore, the choice of H ₂ O ₂ concentration should be made with the goal of achieving efficient pollutant removal while minimizing the environmental impact associated with excessive reagent use.	21
10.	It is preferred to add the general chemical reactions that proceeded throughout each technology. Please revise.		
11.	Line 455: "The rate constant of sulfate radicals to BPA was $k_{SO_4^{\bullet-} BPA} = 1.37 \pm 0.15 \times 10^9 \text{ M}^{-1} \text{ s}^{-1}$." For which order is this rate constant? Please recheck the units.	The rate constant provided in the statement "The rate constant of sulfate radicals to BPA was $k_{SO_4^{\bullet-} BPA} = 1.37 \pm 0.15 \times 10^9 \text{ M}^{-1} \text{ s}^{-1}$ " is a second-order rate constant. The correct units for this rate constant are indeed $\text{M}^{-1} \text{ s}^{-1}$. This unit indicates that it's a second-order rate constant because it has dimensions of concentration ⁻¹ time ⁻¹ .	26